

# NEW DIFFRACTION DATA

## Synthesis and X-ray diffraction data for dibromo-dioxo-(1,10-phenanthroline-N,N')-molybdenum(VI) ( $C_{12}H_8N_2MoBr_2O_2$ )

H. A. Camargo,<sup>1,a)</sup> J. A. Henao,<sup>2</sup> and N. J. Castellanos<sup>3</sup>

<sup>1</sup>Grupo de Investigación en Nuevos Materiales y Energías Alternativas (GINMEA), Universidad Santo Tomás, Facultad de Química Ambiental, Campus Universitario Floridablanca, Santander, Colombia

<sup>2</sup>Grupo de Investigación en Química Estructural (GIQUE). Escuela de Química, Facultad de Ciencias, Universidad Industrial de Santander, A.A. 678, Carrera 27, Calle 9 Ciudadela Universitaria, Bucaramanga – Colombia

<sup>3</sup>Estado Sólido y Catálisis Ambiental (ESCA), Departamento de Química, Facultad de Ciencias, Universidad Nacional de Colombia, Carrera 30 No. 45-03, Bogotá, Colombia

(Received 1 April 2015; accepted 24 November 2015)

The dibromo-dioxo-(1,10-phenanthroline-N,N')-molybdenum(VI) complex ( $C_{12}H_8N_2MoBr_2O_2$ ) was prepared from molybdic acid and 1,10-phenanthroline using hydrobromic acid as solvent. The molybdenum complex solid was separated by filtration and washed with ethyl ether. The X-ray powder diffraction pattern for the title compound was analyzed and found to crystallizes in monoclinic system, space group  $P2_1/c$  (No14) with refined unit-cell parameters  $a = 12.036$  (1),  $b = 9.819$  (1),  $c = 12.671$  (2) Å, and  $\beta = 110.44^\circ$  (1). The volume of the unit cell is  $V = 1403.2$  (3) Å<sup>3</sup>. © 2016 International Centre for Diffraction Data. [doi:10.1017/S0885715615000949]

Key words: dioxomolybdenum complex, X-ray powder diffraction, oxygen atom transfer

### I. INTRODUCTION

1,10-Phenanthroline represents one of the most representative bidentate chelating ligands in coordination chemistry (Accorsi *et al.*, 2009). Its versatility, manifested in the ability to be functionalized in different positions has expanded its coordination compounds in new applications in areas such as medicine (Ganeshpandian *et al.*, 2014; Inci *et al.*, 2015), electronics devices (Kurtz *et al.*, 2015), sensors (Kaur and Alreja, 2015), and new materials to oxidation (Yang *et al.*, 2015) or degradation of organic compounds (Abolhosseini *et al.*, 2014). Structurally, the main difference with the common 2,2'-bipyridine ligand is imposed by the rigid center ring in which the two nitrogen atoms are always held in juxtaposition, increasing the formation rates of bidentate coordination complex (Sammes and Yahioglu, 1994). Among the innumerable metal complexes with 1,10-phenanthroline ligand and phenanthroline derivates ligand, ruthenium (II) (Heidary *et al.*, 2014), copper (II) (Liu *et al.*, 2014), nickel (Cai *et al.*, 2013), and iridium (Kwon *et al.*, 2015) metals are certainly the most widely used. We have reported, over the years, the synthesis of several molybdenum complexes with bidentate chelating ligands such as bipyridyl (Páez *et al.*, 2008, 2009) and bis-pyrazolyl (Castellanos *et al.*, 2012) ligands and studied, under homogeneous and heterogeneous conditions, their properties as oxygen atom transfer agents (Arzoumanian *et al.*, 2010). In a previous work, we have reported the powder diffraction data of a related molybdenum complex (Camargo *et al.*, 2014). Currently, our interest focuses on the study of the electron density effect of different bidentate ligands in oxygen atom transfer processes catalyzed by dioxo-molybdenum entity. In this work, we report the synthesis, molecular

characterization [Fourier transform infrared (FTIR), nuclear magnetic resonance (NMR)], and X-ray powder diffraction (XRPD) data for the compound dibromo-dioxo-(1,10-phenanthroline-N,N')-molybdenum(VI) newly synthesized (Figure 1).

### II. EXPERIMENTAL

#### A. Synthesis

All materials were commercial and were used without further purification unless otherwise noted. All solvents were thoroughly degassed prior to use. Acetonitrile was distilled and kept under argon. Molybdic acid ( $H_2MoO_4 \cdot H_2O$ , 5 g) was dissolved in the minimum volume (ca. 25 ml) of warm concentrated (46%) hydrobromic acid and the mixture was slowly stirred at 25 °C for 1 h. 1,10-Phenanthroline monohydrate ( $C_{12}H_8N_2 \cdot H_2O$ , 5, 50 g) was added in the solution to obtain a yellow solid. The product was separated by filtration, washed with diethyl ether (3 × 100 ml) and recrystallized from acetone (yield 82%). The molecular characterization was carried out with ultraviolet-visible (UV–Vis) spectroscopy which showed two absorption bands at 270 and 350 nm. Infrared spectrometry showed stretching vibrations;  $\nu$ : 1650 (C=N); 1434 (C=C); 940, 895 (Mo=O); and NMR on protons (<sup>1</sup>H NMR, 400 MHz, dimethyl sulfoxide (DMSO)) showed  $\delta$  (ppm) = 8.15 (d-d, 2H), 7.51 (d-d, 2H), 7.02 (s, 2H), 6.82 (d-d, 2H).

#### B. Powder data collection

A small portion of the title compound was gently ground in an agata mortar and sieved to a grain size <38 µm. The specimen was mounted on a polymethyl methacrylate specimen holder. The XRPD pattern was recorded with a D8 ADVANCE BRUKER diffractometer operating in DaVinci

<sup>a)</sup>Author to whom correspondence should be addressed. Electronic mail: hernando.camargo@mail.ustabuca.edu.co

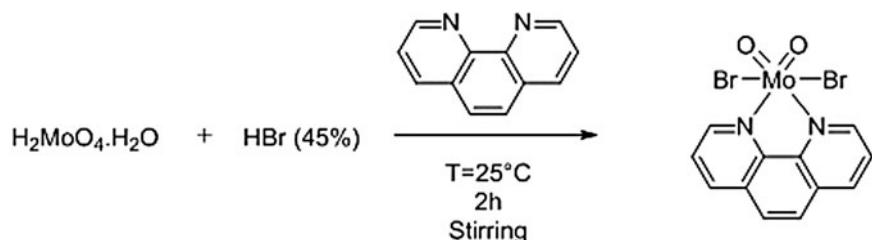


Figure 1. Synthesis of dibromo-dioxo-(1,10-phenanthroline-N,N')-molybdenum(VI).

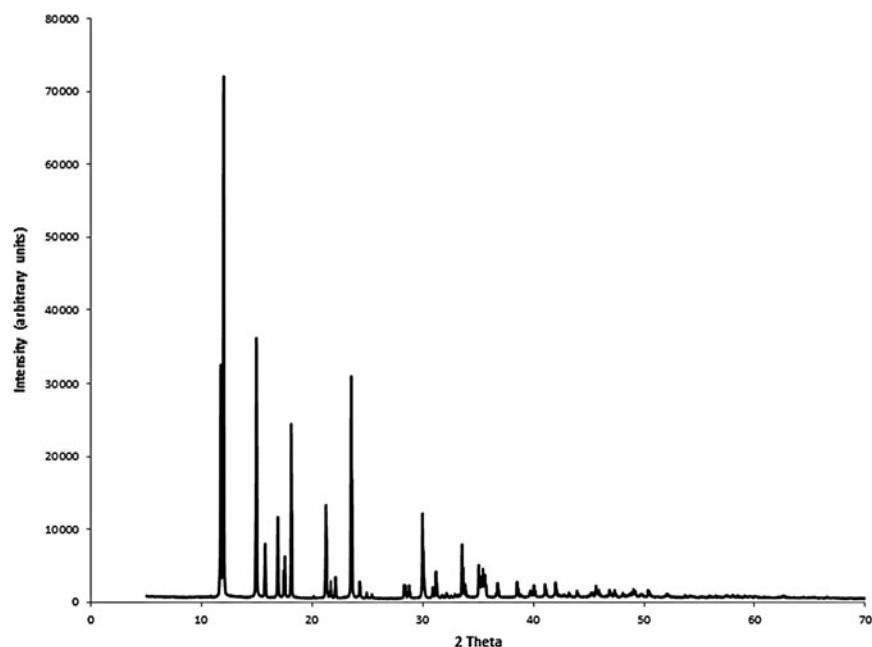


Figure 2. Powder X-ray diffraction pattern of dibromo-dioxo-(1,10-phenanthroline-N,N')-molybdenum(VI).

TABLE I. X-ray Powder Diffraction data of dibromo-dioxo-(1,10-phenanthroline-N,N')-molybdenum(VI).

$2\theta_{\text{obs}}$ ( $^{\circ}$ )	$d_{\text{obs}}$ ( $\text{\AA}$ )	$(I/I_0)_{\text{obs}}$	$H$	$k$	$l$	$2\theta_{\text{cal}}$ ( $^{\circ}$ )	$d_{\text{cal}}$ ( $\text{\AA}$ )	$\Delta 2\theta$ ( $^{\circ}$ )	
11.682	7.5692	46		0	1	11.686	7.5668	0.004	
11.942	7.4049	100		-1	1	0	11.941	7.4055	-0.001
14.906	5.9385	56		0	0	2	14.910	5.9368	0.004
15.696	5.6413	12		2	0	0	15.703	5.6390	0.007
16.850	5.2575	17		-1	1	2	16.849	5.2577	-0.001
17.385	5.0969	6		-2	1	1	17.382	5.0977	-0.003
				0	1	2	{17.442	5.0804	
17.487	5.0674	9		-2	0	2	{17.490	5.0665	0.003
18.053	4.9098	36		0	2	0	18.054	4.9095	0.001
21.197	4.1881	22		1	1	2	21.202	4.1872	0.005
21.629	4.1054	4		2	1	1	21.632	4.1049	0.003
22.058	4.0265	4		1	2	1	22.055	4.0272	-0.003
23.492	3.7839	51		0	2	2	23.495	3.7834	0.003
23.670	3.7558	1		3	0	0	23.648	3.7593	-0.022
24.240	3.6688	4		-2	1	3	24.251	3.6671	0.011
24.876	3.5764	1		-3	1	2	24.880	3.5758	0.004
25.345	3.5113	1		3	1	0	25.349	3.5108	0.004
28.263	3.1551	3		0	3	1	{28.261	3.1553	-0.002
				-1	0	4	{28.267	3.1546	
28.371	3.1433	3		1	3	0	28.371	3.1433	0.000
28.719	3.106	3		-3	2	1	{28.724	3.1055	0.005
				3	1	1	{28.749	3.1028	
				-3	2	0	{29.912	2.9848	
29.934	2.9826	20		-4	0	2	{29.938	2.9822	0.004
30.059	2.9705	4		0	0	4	{30.081	2.9684	0.022
				1	3	1	{30.085	2.9680	
				-2	1	4	{30.093	2.9672	

Continued

TABLE I. Continued

$2\theta_{\text{obs}}$ ( $^{\circ}$ )	$d_{\text{obs}}$ ( $\text{\AA}$ )	$(I/I_0)_{\text{obs}}$	$H$	$k$	$l$	$2\theta_{\text{cal}}$ ( $^{\circ}$ )	$d_{\text{cal}}$ ( $\text{\AA}$ )	$\Delta 2\theta$ ( $^{\circ}$ )
30.847	2.8964	3	-1	3	2	30.840	2.8970	-0.007
31.149	2.869	7	-2	3	1	31.145	2.8694	-0.004
			-3	0	4	31.171	2.8670	
			0	3	2	31.179	2.8663	
31.701	2.8203	1	4	0	0	31.710	2.8195	0.009
32.102	2.786	1	1	2	3	32.101	2.7860	-0.001
			-3	2	3	32.141	2.7827	
32.875	2.7222	1	3	2	1	32.880	2.7218	0.005
33.177	2.6981	1	-4	1	3	33.183	2.6976	0.006
33.509	2.6721	13	1	3	2	33.505	2.6725	-0.004
33.617	2.6638	2	3	1	2	33.601	2.6650	-0.016
33.791	2.6505	3	2	3	1	33.788	2.6507	-0.003
35.021	2.5602	8	1	1	4	35.023	2.5600	0.002
			-4	2	1	35.133	2.5523	
35.178	2.5491	4	-4	2	2	35.182	2.5488	0.004
35.295	2.5409	5	0	2	4	35.305	2.5402	0.010
			-3	3	1	35.373	2.5355	
35.402	2.5335	6	-4	0	4	35.405	2.5332	0.003
			0	3	3	35.564	2.5223	
35.577	2.5214	6	-2	3	3	35.582	2.5210	0.005
36.728	2.4450	3	4	2	0	36.728	2.4450	0.000
38.492	2.3369	4	-5	1	2	38.495	2.3367	0.003
38.819	2.318	1	1	4	1	38.820	2.3179	0.001
			-1	3	4	39.649	2.2713	
			-2	4	1	39.673	2.2700	
39.690	2.2691	2	0	4	2	39.701	2.2685	0.011
			2	1	4	39.988	2.2528	
40.018	2.2512	3	-4	2	4	40.018	2.2512	0.000
			-2	4	0	40.027	2.2507	
			0	3	4	41.015	2.1988	
41.020	2.1985	3	-5	1	0	41.024	2.1983	0.004
41.962	2.1513	4	-5	1	4	41.965	2.1512	0.003
			0	2	5	42.241	2.1378	
42.268	2.1365	1	-5	2	1	42.265	2.1366	-0.003
			-4	3	0	42.274	2.1362	
			-5	2	3	42.727	2.1146	
42.734	2.1142	1	3	3	2	42.740	2.1139	0.006
			-2	0	6	42.783	2.1119	
43.180	2.0934	1	-3	4	1	43.177	2.0935	-0.003
43.905	2.0605	2	1	3	4	43.908	2.0604	0.003
45.194	2.0047	1	-6	0	2	45.193	2.0047	-0.001
			-4	3	4	45.228	2.0033	
45.619	1.9870	3	1	4	3	45.621	1.9869	0.002
			-3	4	3	45.650	1.9857	
45.885	1.9761	2	-4	0	6	45.874	1.9766	-0.011
			-5	3	2	46.828	1.9385	
46.842	1.9379	2	-4	1	6	46.848	1.9377	0.006
			0	5	1	46.854	1.9375	
			-1	4	4	46.858	1.9373	
			-5	3	1	47.274	1.9212	
47.303	1.9201	2	5	2	1	47.301	1.9202	-0.002
			2	1	5	47.309	1.9199	
			-6	0	4	47.313	1.9197	
48.056	1.8918	1	0	4	4	48.058	1.8917	0.002
			1	5	1	48.061	1.8916	
			-5	2	5	48.561	1.8733	
48.572	1.8729	1	-1	5	2	48.574	1.8728	0.002
48.778	1.8655	1	-2	5	1	48.783	1.8653	0.005
			5	1	2	48.786	1.8652	
			-3	4	4	48.800	1.8646	
			0	5	2	48.806	1.8644	
49.012	1.8571	2	-5	3	0	49.007	1.8573	-0.005
49.182	1.8511	1	-4	4	0	49.173	1.8514	-0.009
			-4	3	5	49.210	1.8501	
			0	2	6	49.629	1.8354	
49.668	1.8341	1	-4	2	6	49.684	1.8335	0.016

*Continued*

TABLE I. Continued

$2\theta_{\text{obs}}$ ( $^{\circ}$ )	$d_{\text{obs}}$ ( $\text{\AA}$ )	$(I/I_0)_{\text{obs}}$	$H$	$k$	$l$	$2\theta_{\text{cal}}$ ( $^{\circ}$ )	$d_{\text{cal}}$ ( $\text{\AA}$ )	$\Delta 2\theta$ ( $^{\circ}$ )
50.337	1.8113	2	1	3	5	50.320	1.8119	
52.067	1.7551	1	1	6	6	50.343	1.8111	0.006
			6	2	0	52.057	1.7554	-0.010
				-1	4	52.103	1.7539	
				0	4	53.657	1.7068	
53.703	1.7054	1	-5	4	1	53.677	1.7062	-0.026
				7	1	57.960	1.5899	
57.982	1.5893	1	5	4	1	57.965	1.5897	-0.017
				-6	3	58.090	1.4842	
				-7	2	58.096	1.4840	
62.572	1.4833	1	-4	2	8	62.559	1.4836	-0.013

TABLE II. Parameters obtained by XRPD for the compound dibromo-dioxo-(1,10-phenanthroline-N,N')-molybdenum(VI).

Dibromo-dioxo-(1,10-phenanthroline-N,N')-molybdenum(VI)	
$a$ ( $\text{\AA}$ )	12.036 (1)
$b$ ( $\text{\AA}$ )	9.819 (1)
$c$ ( $\text{\AA}$ )	12.671 (2)
$\beta$ ( $^{\circ}$ )	110.44 (1)
$V$ ( $\text{\AA}^3$ )	1403.2 (2)
$Z$	4
$M_{20}$	41.3
$F_{30}$	65.8 (0.0056, 82)
$D_m$ ( $\text{g cm}^{-3}$ )	2.19

geometry equipped with an X-ray tube ( $\text{CuK}\alpha$  radiation;  $\lambda = 1.5406 \text{\AA}$ , 40 kV, and 30 mA) using a nickel filter and a one-dimensional LynxEye detector. A receiving slit (RS) of 0.6 mm and primary and secondary soller slits (SS) of  $2.5^{\circ}$  were used. The scan range was  $5^{\circ}$ – $70^{\circ}2\theta$  with a step size of  $0.01526^{\circ}$  and a count time of 2 s per step. Powder data were collected at room temperature (298 K).

Powder analytical software was used to remove the background (Sonneveld and Visser, 1975), to smooth the pattern (Savitzky and Golay, 1964), and to eliminate the  $\text{K}\alpha_2$  component from each reflection (Rachinger, 1948). The second derivative method was used to determine the position and intensities of the diffraction maxima from each reflection.

### III. RESULTS AND DISCUSSION

The X-ray powder pattern of the compound dibromo-dioxo-(1,10-phenanthroline-N,N')-molybdenum(VI) is shown in Figure 2. The XRPD data for the compound are given in Table I. All reflections were indexed successfully using the DICVOL06 program (Boultif and Louér, 2004) in a monoclinic unit cell. An absolute error of  $0.03^{\circ}$  ( $2\theta$ ), was assigned to all the peak positions used in the indexing of the pattern. The space group,  $P2_1/c$  (No. 14), estimated by the program CHEKCELL (Laugier and Bochu, 2002) was compatible with the systematic absences and with the crystal density. The lattice parameters of the compound (2) were refined with the program NBS\*AIDS83 software (Miguell *et al.*, 1981). Its crystal data, X-ray density and figures of merit  $M_{20}$  (de Wolff, 1968) and  $F_{20}$  (Smith and Snyder, 1979) of the compound under study are compiled in Table II.

### SUPPLEMENTARY MATERIAL

To view supplementary material for this article, please visit <http://dx.doi.org/10.1017/S0885715615000949>.

### ACKNOWLEDGEMENTS

The authors thank Centro de Investigaciones de Universidad Santo Tomás (Bucaramanga-Colombia) for their support with the project approved in the VII internal call of research projects and the Laboratorio de Rayos-X PTG, Vicerrectoría de Investigación y Extensión of Universidad Industrial de Santander (Bucaramanga-Colombia) for data collection.

- Abolhosseini, Sh. A., Mahjoub, A. R., Eslami-Moghadam, M., and Fakhri, H. (2014). "Dichloro (1,10-phenanthroline-5,6-dione) palladium (II) complex supported by mesoporous silica SBA-15 as a photocatalyst for degradation of 2,4-dichlorophenol," *J. Mol. Struct.* **1076**, 568–575.
- Accorsi, G., Listorti, A., Yoosaf, K., and Armaroli, N. (2009). "1,10-Phenanthrolines: versatile building blocks for luminescent molecules, materials and metal complexes," *Chem. Soc. Rev.* **38**, 1690–1700.
- Arzoumanian, H., Castellanos, N. J., Martínez, F. O., Páez-Mozo, E. A., and Ziarelli, F. (2010). "Silicon-assisted direct covalent grafting on metal oxide surfaces: synthesis and characterization of carboxylate N, N'-ligands on  $\text{TiO}_2$ ," *Eur. J. Inorg. Chem.* **2010**, 1633–1641.
- Boultif, A. and Louér, D. (2004). "Indexing of powder diffraction patterns of low symmetry lattices by successive dichotomy method," *J. Appl. Crystallogr.* **37**, 724–731.
- Cai, Z., Liu, L., and Zhou, M. (2013). "Synthesis of nickel(II) complexes containing modified phenanthroline ligands for potential nonlinear optical applications," *Opt. Mater.* **35**, 1481–1486.
- Camargo, H. A., Castellanos, N. J., Rosas, C. C., and Henao, J. A. (2014). "Synthesis and X-ray diffraction data of dichlorodioxido (4,4-dimethoxycarbonyl-2,2'-bipyridyl) molybdenum(VI)," *Powder Difffr.* **29**, 1, 42–45.
- Castellanos, N., Martínez, F., Páez-Mozo, E., Ziarelli, F., and Arzoumanian, H. (2012). "Bis(3,5-dimethylpyrazol-1-yl)acetate bound to titania and complexed to molybdenum dioxido as a bidentate N,N'-ligand. Direct comparison with a bipyridyl analog in a photocatalytic arylalkane oxidation by  $\text{O}_2$ ," *Transit. Met. Chem.* **37**, 629–637.
- de Wolff, P. M. (1968). "A simplified criterion for the reliability of a powder pattern indexing," *J. Appl. Crystallogr.* **1**, 108–113.
- Ganeshpandian, M., Ramakrishnan, S., Palaniandavar, M., Suresh, E., Riyasdeen, A., and Akbarsha, M. A. (2014). "Mixed ligand copper(II) complexes of 2,9-dimethyl-1,10-phenanthroline: tridentate 3N primary ligands determine DNA binding and cleavage and cytotoxicity," *J. Inorg. Biochem.* **140**, 202–212.
- Heidary, D. K., Howerton, B. S., and Glazer, E. C. (2014). "Coordination of hydroxyquinolines to a ruthenium bis-dimethyl-phenanthroline scaffold

- radically improves potency for potential as antineoplastic agents," *J. Med. Chem.* **57**, 8936–8946.
- İnci, D., Aydin, R., Yılmaz, D., Gençkal, H. M., Vatan, Ö., Çinkılıç, N., and Zorlu, Y. (2015). "New water-soluble copper (II) complexes including 4,7-dimethyl-1,10-phenanthroline and l-tyrosine: synthesis, characterization, DNA interactions and cytotoxicities," *Spectrochim. Acta A: Mol. Biomol. Spectrosc.* **136B**, 761–770.
- Kaur, N. and Alreja, P. A. (2015). "Novel 1,10-phenanthroline based chemosensor for differential metal ion sensing and constructing molecular logic gates," *Tetrahedron Lett.* **56**, 182–186.
- Kurtz, D. A., Dhakal, B., Hulme, R. J., Nichol, G. S., and Felton, G. A. N. (2015). "Correlations between photophysical and electrochemical properties for a series of new Mn carbonyl complexes containing substituted phenanthroline ligands," *Inorg. Chim. Acta* **427**, 22–26.
- Kwon, Y., Sunesh, C. D., and Choe, Y. (2015). "Light-emitting properties of cationic iridium complexes containing phenanthroline based ancillary ligand with blue-green and green emission colors," *Opt. Mater.* **39**, 40–45.
- Laugier, J. and Bochu, B. (2002). *CHEKCELL*. "LMGP-Suite Suite of Programs for the interpretation of X-ray Experiments", ENSP/Laboratoire des Matériaux et du Génie Physique, BP 46. 38042 Saint Martin d'Hères, France. <http://www.inpg.fr/LMGP> and <http://www CCP14.ac.uk/tutorial/lmp/>
- Liu, B., Pan, S., Liu, B., and Chen, W. (2014). "Di-, tri-, and tetrานuclear copper(I) complexes of phenanthroline-linked dicarbene ligands," *Inorg. Chem.* **53**, 10485–10497.
- Miguell, A. D., Hubberd, C. R., and Stalick, J. K. (1981). "NBS\* AIDS80: a FORTRAN program for crystallographic data evaluation," National Bureau of Standards (USA), Tech. Note 1141.
- Páez, C. A., Castellanos, N. J., Martínez, O. F., Ziarelli, F., Agrifoglio, G., Páez-Mozo, E. A., and Arzoumanian, H. (2008). "Oxygen atom transfer photocatalyzed by molybdenum(VI) dioxodibromo-(4,4'-dicarboxylato-2,2'-bipyridine) anchored on TiO<sub>2</sub>," *Catal. Today* **133**–135, 619–624.
- Páez, C. A., Lozada, O., Castellanos, N. J., Martínez, F. O., Ziarelli, F., Agrifoglio, G., Páez-Mozo, E. A., and Arzoumanian, H. (2009). "Arylalkane photo-oxidation under visible light and O<sub>2</sub> catalyzed by molybdenum(VI)dioxo-dibromo (4,4'-dicarboxylato-2,2'-bipyridine) anchored on TiO<sub>2</sub>," *J. Mol. Catal. A: Chem.* **299**, 53–59.
- Rachinger, W. A. (1948). "A correction for the  $\alpha_1 \alpha_2$  doublet in the measurement of widths of X-ray diffraction lines," *J. Sci. Instrum.* **25**, 254.
- Sammes, P. G. and Yahioglu, G. (1994). "1,10-Phenanthroline: a versatile ligand," *Chem. Soc. Rev.* **23**, 327–334.
- Savitzky, A. and Golay, M. J. (1964). "Smoothing and differentiation of data by simplified least squares procedures," *Anal. Chem.* **36**, 1627–1639.
- Smith, G. S. and Snyder, R. L. (1979). " $F_N$ : a criterion for rating powder diffraction patterns and evaluating the reliability of powder-pattern indexing," *J. Appl. Crystallogr.* **12**, 60–65.
- Sonneveld, E. J. and Visser, J. W. (1975). "Automatic collection of powder diffraction data from photographs," *J. Appl. Crystallogr.* **8**, 1–7.
- Yang, Y., Wang, W., Liu, Y., Wang, F., Chai, D., and Lei, Z. (2015). "Pd nanoparticles supported on phenanthroline modified carbon as high active electrocatalyst for ethylene glycol oxidation," *Electrochim. Acta* **154**, 1–8.